



ATTACHMENT A

Clean Replacement Paragraphs

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At the following location, replace the previously provided paragraph with the following clean paragraph.

Page 9, line 13. - pg. 10, line 2

Scanned

H(CF₂)₆CH₂OH (664.00 g, 2.00 moles), 55% nitric acid (114.55 g, 1.00 mole) and FeCl₂ · nH₂O (0.0066 g) were placed into an autoclave equipped with a 1000-ml glass pressure vessel, a fluoroplastic upper cover, stirring blades, a thermometer protection tube, a fluoroplastic insert tube, a pressure gauge, a safety valve and a supply line from an oxygen bomb. The mixture was stirred with heating, upon which reaction pressure began to increase. 3.1 hours after the start of heating, the reaction temperature rose to 125 and the reaction pressure increased to 0.6 MPa (gauge pressure; the same hereinafter). From that time, oxygen was fed into the gas phase at various times in an amount of 0.35 g (11.00 mmoles) per oxygen fed process, whereby the reaction pressure was controlled to 0.6 MPa. 6.5 hours after the start of heating, complete consumption of the starting fluoroalkyl alcohol was confirmed by gas chromatography to confirm the completion of the reaction [H(CF₂)₆CH₂OH conversion: 100.0 g.c.% (gas chromatography %; the same hereinafter); H(CF₂)₆COOH selectivity: 100.0 g.c.%]. In total, 46.72 g (1.46 moles) of oxygen was fed into the gas phase by the end of the reaction. After completion of the reaction, oxygen was continuously supplied so as to convert residual nitrogen oxides into nitric acid. Then the residual pressure was released. Because of the reaction mixture being provided in the form of two layers of

liquids, 765.78 g of a crude carboxylic acid $[H(CF_2)_6COOH]$ was obtained in a concentrated form from the lower layer by means of liquid-liquid separation at the cease of stirring. The crude carboxylic acid was purified by distillation under reduced pressure, giving 499.75 g of a carboxylic acid $[H(CF_2)_6COOH]$ in high purity (99 g.c.% or higher) with an isolation yield of 65.26 mole %.

B)
conclude

